

図3 発現強度のスカタープロットと発現強度に応じた段階的遺伝子選択法

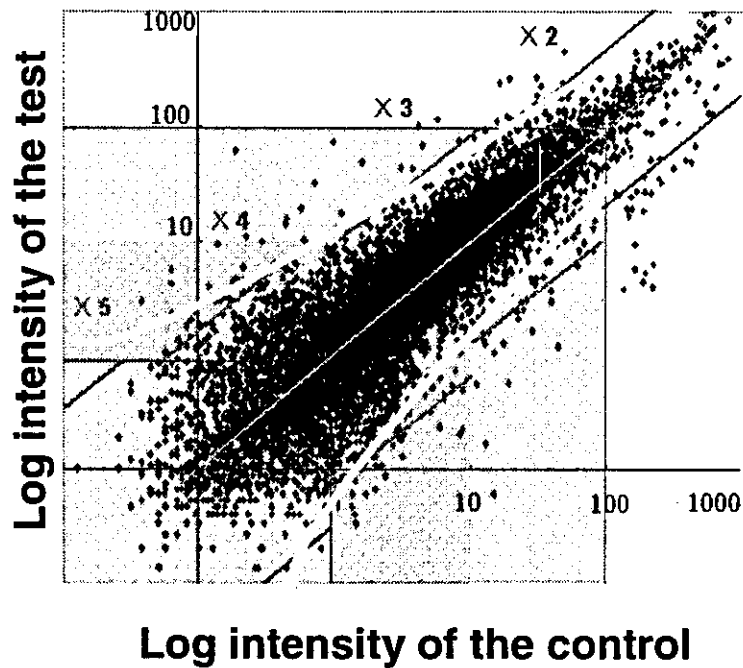


図4 遺伝子発現データベースの構造

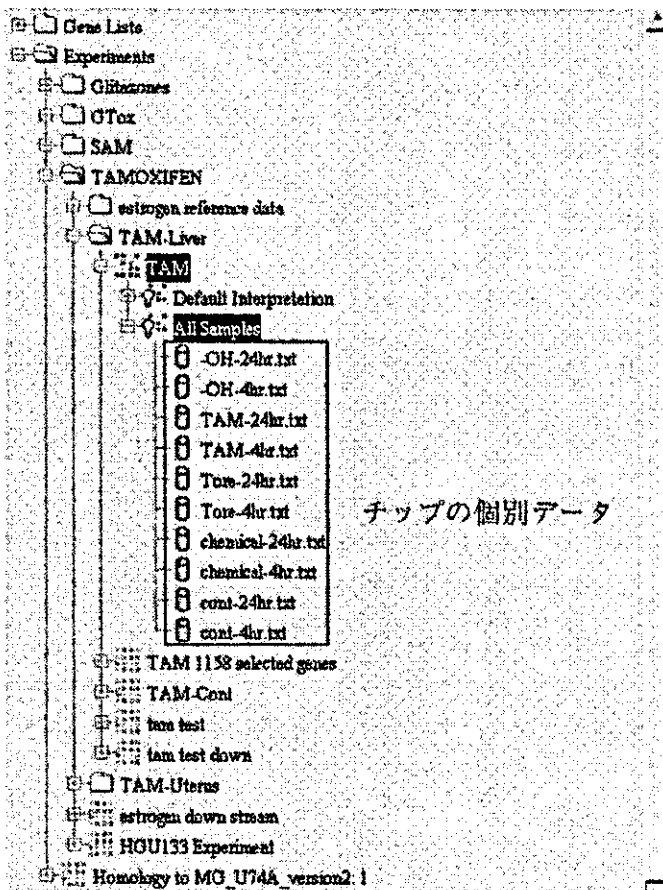


図5 ヒトデータとマウスデータの直接比較法 (HU133A から MOE430 への変換の例)

1. ExcelにてHU133A MOE430 アレイのプローブ ID をリストアップした表を作成。

	A	B
1	221517_s_ε1451719_at	
2	209708_at	1422643_at
3	221407_at	1423019_at
4	214886_s_ε1417707_at	
5	204843_s_ε1452915_at	
6	202780_at	1428140_at

2. このファイルを MOE430.homology という名前で、GeneSpring のHU133 チップフォルダ内にある "Homology Tables" というフォルダへ保存する。

3. これにより、自動的にHU133A チップでの遺伝子リストをマウスのゲノムへ変換するための準備が整うが、GeneSpring 上で変換操作を行うと、両者のデータは横並びには表示されず、対応するマウスのチップ上のプローブ ID のみが右のように表示される。

Summary for 133A

Total number of genes 22,921

1: Systems	2: Common Na	23: mouse homolog	27: A
224	202077_at	1428160_at	
224	202078_at	1416678_at	
224	202079_s_ε1451042		
224	202080_s_ε1451042	1428327_at	
224	202081_at	1415442_at	
224	202082_at	1451707_at	

Buttons: Add Column, Add Gene, Import from File, Delete Column, Delete Gene, OK, Cancel, Help

4. 次に mouse homolog ID を含む HU133A データを表形式で Excel 上へ書き出すと右の表の左側の赤色で示した部分に mouse homolog (青色) が追加される。

Gene	TAIR ID	TAIR ID	TAIR ID	Gene	Mouse homolog
224	202077_at	1428160_at			
224	202078_at	1416678_at			
224	202079_s_ε1451042				
224	202080_s_ε1451042	1428327_at			
224	202081_at	1415442_at			
224	202082_at	1451707_at			

5. 次に、この mouse homolog を用いて MOE430 用の遺伝子リストを作成し、GeneSpring 上にて対応するマウスチップのデータから対応するデータをリストアップし、同様に Excel 上へ転送する。

Gene	TAIR ID	TAIR ID	TAIR ID	Gene	Mouse homolog
224	202077_at	1428160_at			
224	202078_at	1416678_at			
224	202079_s_ε1451042				
224	202080_s_ε1451042	1428327_at			
224	202081_at	1415442_at			
224	202082_at	1451707_at			

6. Mouse homolog の項をキーとして HU133A および MOE430 のデータを並び替え、一部重複する項目を手作業にて整理することにより両者を一つの Excel ファイル上で横並びの表として結合させる。これにより、両者のデータを用いてスクリーンプロットを書かせるなどの直接比較が可能となる。

図6 タモキシフェン処理ヒト肝初代培養細胞でのデータとマウス肝臓でのデータの直接比較

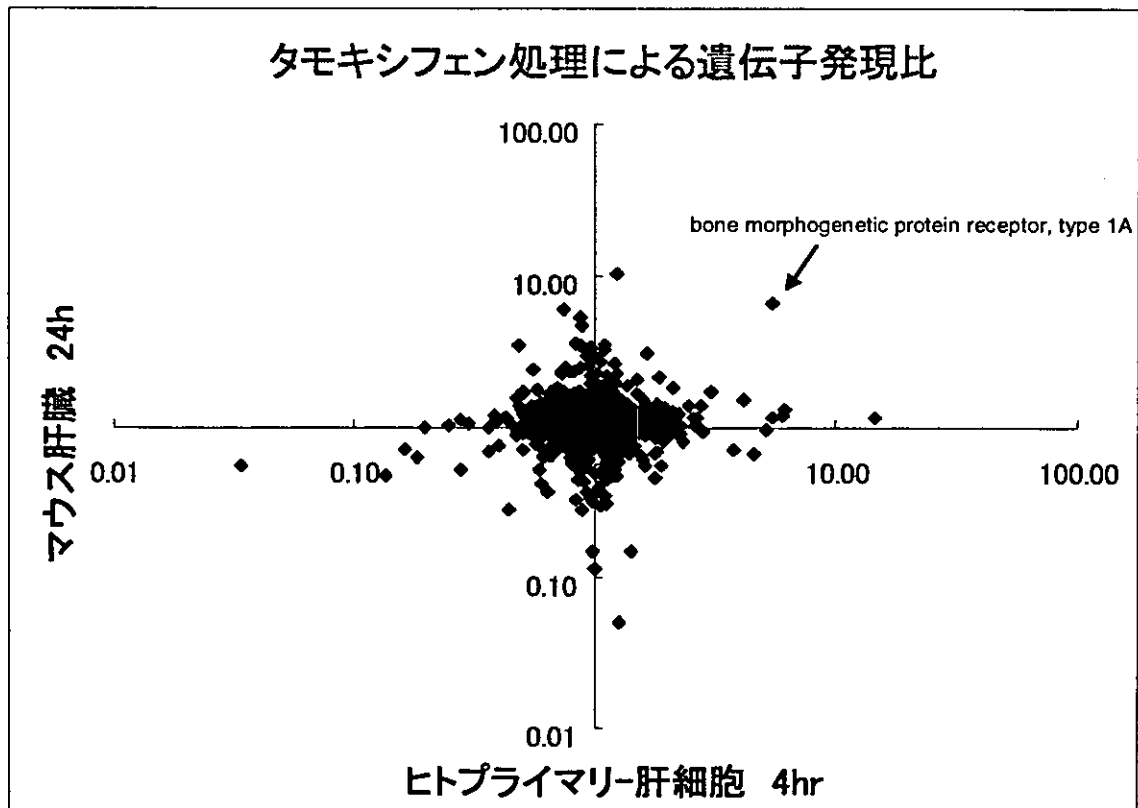


表19 グリタゾン化合物処理による in vivo(マウス)と in vitro(ヒト)データの共通性の比較

Mouse in vivo																				Human in vitro										Description								
FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP20	FP	FP	FP	FP	FP	FP	FP	FP	FP	FP		FP	FP	FP	FP	FP	FP	FP	FP
1.75	1.75	1.52	1.70	1.12	1.40	1.33	1.06	0.93	0.65	0.78	1.14	0.48	1.59	0.74	5.72	4.14	1.87	1.65	1.04	1.73	2.26	1.50	0.82	0.90	0.84	1.10	1.14	1.04	1.35	1.57	2.88	1.91	1.40	1.58	Spin 1			
3.05	3.00	1.93	2.59	2.82	2.80	4.41	4.18	2.35	2.34	4.36	1.88	1.14	1.55	0.80	1.73	1.23	1.31	1.55	2.74	0.98	1.19	2.88	1.01	5.63	4.11	4.31	1.00	0.85	1.34	1.31	1.41	0.79	1.93	1.34	1.07	CD36 antigen		
1.74	1.87	1.88	1.81	1.85	1.46	2.31	1.38	2.00	1.90	2.04	1.82	1.08	0.87	1.00	0.97	0.94	1.04	1.39	1.18	1.24	1.13	1.27	1.03	0.85	0.76	0.67	1.40	1.58	1.94	0.83	0.55	1.32	4.56	3.02	3.47	expressed sequence AC15346		
30.35	65.24	75.35	91.04	42.18	45.47	1.46	2.77	3.33	0.42	2.54	1.46	1.78	1.42	0.81	1.85	0.30	1.20	1.27	0.93	1.40	0.85	1.68	0.74	0.85	0.88	0.66	1.14	1.02	1.15	1.56	1.41	2.83	0.97	0.79	1.33	ferroxidase synthase		
3.95	3.98	3.98	3.42	3.92	2.89	1.30	1.11	1.02	0.43	1.00	1.00	0.98	1.13	1.09	0.86	0.91	1.00	1.40	1.10	1.14	0.99	1.07	1.01	2.89	0.93	0.98	1.21	0.16	0.61	7.89	1.94	7.65	0.51	5.39	2.86	cathelin 2		
1.34	1.15	2.16	2.35	1.65	1.86	1.05	0.81	1.54	0.83	1.04	1.33	0.86	0.47	0.88	0.86	1.43	0.64	0.97	0.77	0.94	0.80	0.88	1.43	0.82	1.30	0.84	0.84	1.32	0.83	4.78	6.39	10.92	21.66	32.40	19.94	fatty acid binding protein 4, adipocyte		
3.48	2.67	2.02	2.36	2.20	2.89	1.70	1.61	1.30	1.08	1.24	1.09	1.51	1.48	1.25	1.43	0.88	1.16	1.33	1.55	1.25	1.32	1.40	1.17	1.24	1.35	1.32	1.39	1.41	1.54	2.32	3.79	4.25	2.31	2.35	2.83	PI8D (cytochrome) oxidoreductase		
1.24	1.41	1.15	1.25	1.33	1.17	1.05	1.41	1.24	1.38	1.28	0.99	0.70	0.95	0.93	0.98	0.99	1.12	1.17	1.19	1.23	0.64	1.22	1.18	0.80	1.03	0.90	1.00	0.97	0.97	0.35	4.63	5.97	4.17	8.03	10.09	fatty acid binding protein 1, liver		
0.55	0.37	0.31	0.33	0.88	0.56	1.52	0.45	1.39	1.21	1.51	1.13	0.71	0.98	0.83	0.64	1.17	0.57	0.85	0.84	0.84	1.26	0.71	1.45	0.91	0.67	0.73	0.41	1.61	0.93	0.82	0.13	0.14	0.47	0.25	0.14	acute cancer family 24 member 1		
0.76	0.45	0.78	0.62	0.48	0.64	0.41	0.84	0.74	0.76	1.28	1.39	0.86	0.89	0.88	0.55	1.61	0.86	0.80	0.89	1.43	1.20	1.16	1.43	0.65	0.61	0.72	0.39	0.91	0.98	0.32	0.36	0.28	1.27	0.87	0.64	cardiac respiratory adenylylase protein		
1.19	1.46	1.80	1.98	1.56	1.46	1.24	1.52	1.26	1.48	1.67	1.27	0.89	0.89	1.09	0.80	0.78	1.08	0.77	0.88	1.07	1.05	1.10	0.69	0.52	0.46	0.38	0.53	1.41	0.94	0.17	0.37	0.39	2.00	0.80	0.54	FMN-cDNA 121001E11 gene		

* 各処理において共通性の高かった遺伝子に関して異種データの相同遺伝子のデータを比較した。

ピンクは発現上昇、水色は発現低下

図7 DEN 処理したマウス肝臓の可溶性画分の2次元電気泳動による解析

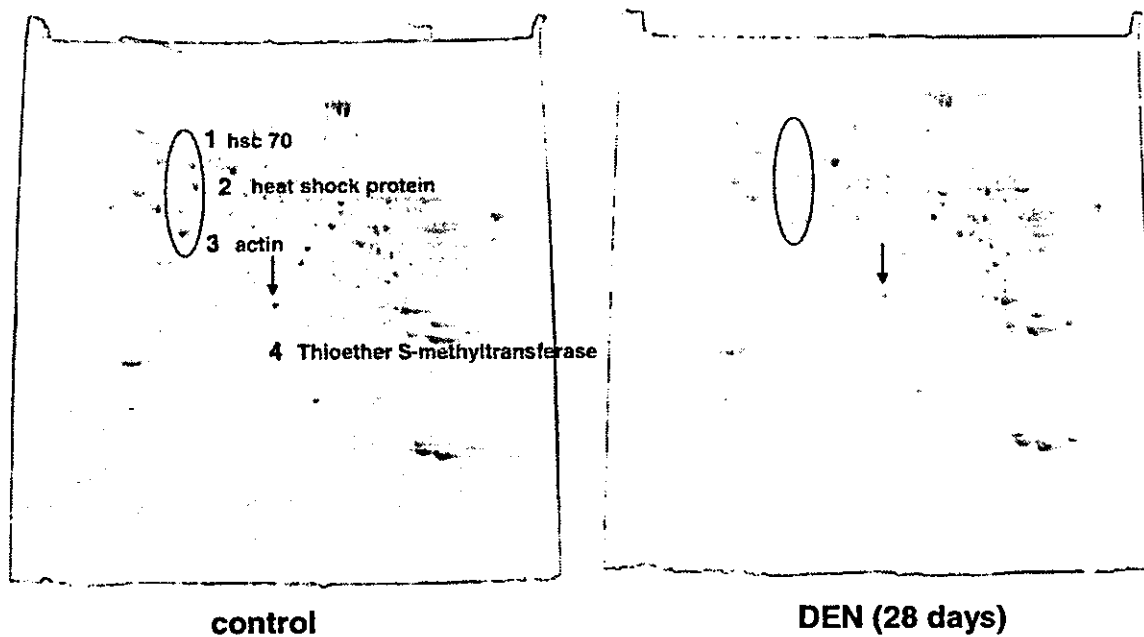


図8 ペプチドマスフィンガープリンティング法によるタンパクの同定例

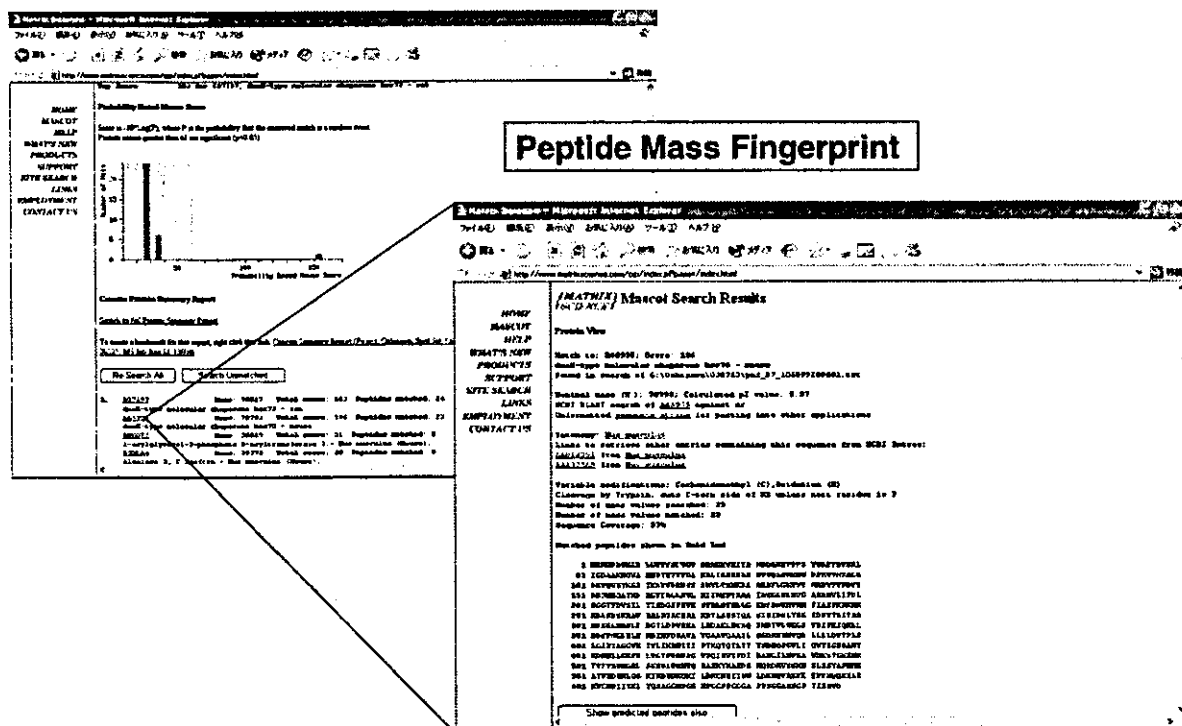


図9 2次元液体クロマトグラムシステム

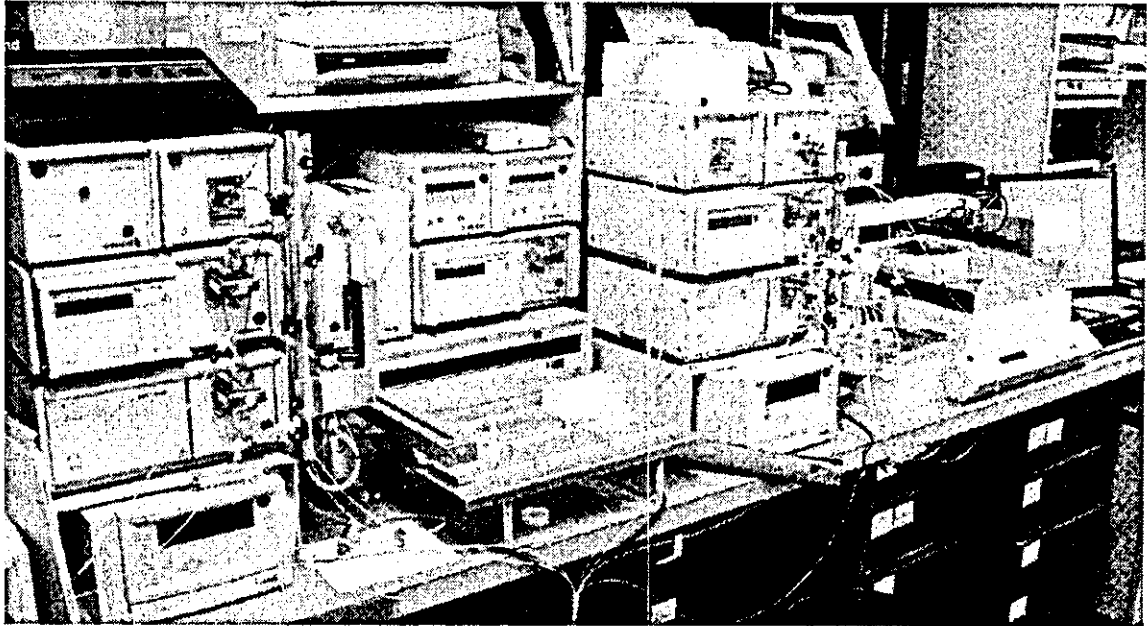
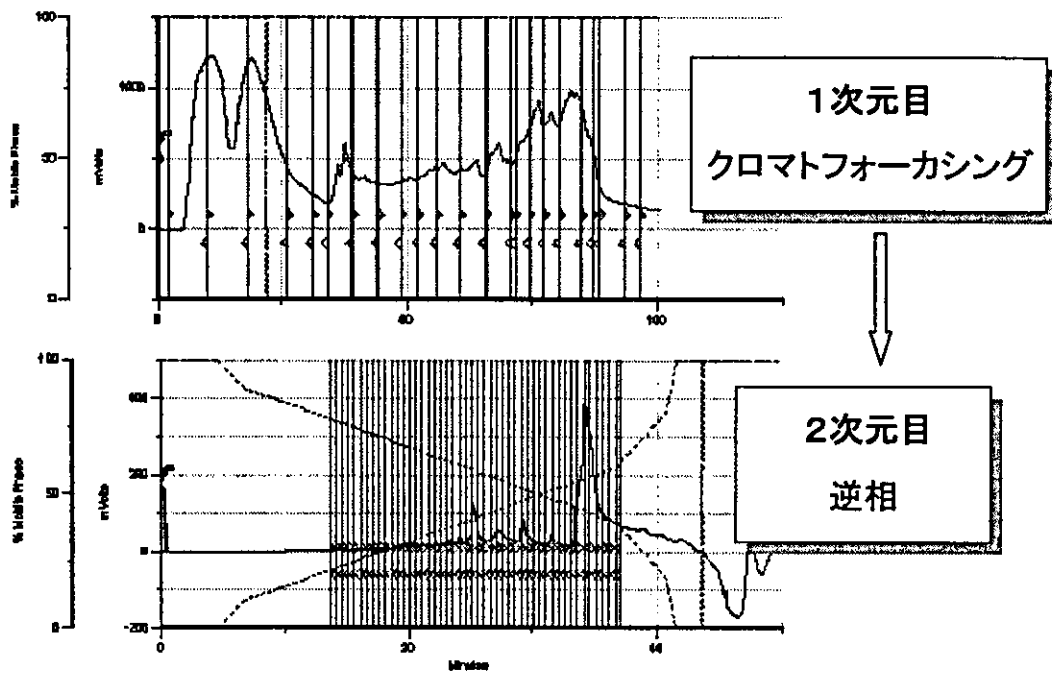


図10 タンパク質溶液の2次元分離例



別添4

II. 研究成果の刊行に関する一覧表

平成14年度～16年度

	発表者氏名	論文タイトル名	発表誌名	巻号	ページ	出版年
1)	Takayoshi Suzuki, Yuki Nagano, Akiyasu Kouketsu, Azusa Matsuura, Sakiko Maruyama, Mineko Kurotaki, Hidehiko Nakagawa, <u>Naoki Miyata</u>	Novel inhibitors of human histone deacetylases: design, synthesis, enzyme inhibition, and cancer cell growth inhibition of SAHA-based non- hydroxamates	<i>J. Med. Chem.</i>	48	1019-1032	2005
2)	Shinya Usui, Takayoshi Suzuki, Yoshifumi Hattori, Kazuma Etoh, Hiroki Fujieda, Makoto Nishizuka, Masayoshi Imagawa, Hidehiko Nakagawa, Kohfuku Kohda, <u>Naoki Miyata</u>	Design, synthesis and biological activity of novel PPAR γ ligands based on rosiglitazone and 15d-PGJ2	<i>Bioorganic & Medicinal Chemistry Letters</i>	15	1545-1551	2005
3)	Takayoshi Suzuki, Azusa Matsuura, Akiyoshi Kouketsu, Hidehiko Nakagawa, <u>Naoki Miyata</u>	Identification of a potent non-hydroxamate histone deacetylase inhibitor by mechanism-based drug design	<i>Bioorganic & Medicinal Chemistry Letters</i>	15	331-335	2005
4)	鈴木孝禎、宮田直樹	ヒストン脱アセチル化酵素阻害剤開発の最前線	ファルマシア	41	<i>in press</i>	2005
5)	鈴木孝禎、中川秀彦、宮田直樹	癌の分子標的治療薬の開発:非ヒドロキサム酸系ヒストン脱アセチル化酵素阻害薬の設計、合成と生物活性評価	有機合成化学協会誌	63	<i>in press</i>	2005
6)	Takayoshi Suzuki, Akiyasu Kouketsu, Azusa Matsuura, Arihiro Kohara, <u>Shin-ichi Ninomiya</u> , Kohfuku Kohda, <u>Naoki Miyata</u>	Thiol-based SAHA analogues as potent histone deacetylase inhibitors,	<i>Bioorganic & Medicinal Chemistry Letters</i>	14(12)	3313-3317	2004
7)	Masaaki Kurihara, Abu Shara Shamsur Rouf, Hisao Kansui, Hiroyuki Kagechika, <u>Haruhiro Okuda</u> , <u>Naoki Miyata</u>	Design and synthesis of cyclic urea compounds: a pharmacological study for retinoidal acty	<i>Bioorganic & Medicinal Chemistry Letters</i>	14(16)	4131-4134	2004
8)	Nobuyuki Sera, Hiroshi Tokiwa, Hideo Utsumi, Shigeki Sasaki, Kiyoshi Fukuhara, <u>Naoki Miyata</u>	Associations between chemical properties and oxidative damage due to nitrophenanthrenes and their related compounds in primary rat hepatocytes	<i>Polycyclic Aromatic Compounds</i>	24(4-5)	487-500	2004
9)	I. Nakanishi, K. Ohkubo, K. Miyazaki, W. Hakamata, S. Urano, T. Ozawa, <u>H. Okuda</u> , S. Fukuzumi, N. Ikota, and K. Fukuhara	A planar catechin analogue having a more negative oxidation potential than (+)-catechin as an electron-transfer antioxidant against a peroxy radical	<i>Chem. Res. Toxicol.</i>	17	26-31	2004
10)	Yamada K, <u>Suzuki T</u> , Kohara A, Hayashi M, Mizutani T, Saeki K	In vivo mutagenicity of benzo[f]quinoline, benzo[h]quinoline, and 1,7-phenanthroline using the lac Z transgenic mice	<i>Mutation Research</i>	559	83-95	2004
11)	Arlt VM, Zhan L, Schmeiser HH, Honma M, Hayashi M, Phillips DH, <u>Suzuki T</u>	DNA adducts and mutagenic specificity of the ubiquitous environmental pollutant 3-nitrobenzanthrone in Muta Mouse	<i>Environ. Mol. Mutagen.</i>	43	186-195	2004
12)	Li Zhana, Hiroko Sakamoto, Mayumi Sakuraba, De-Sheng Wu, Li-Shi Zhang, <u>Takayoshi Suzuki</u> , Makoto Hayashi, Masamitsu Honma	Genotoxicity of micro-LR in human lymphoblastoid TK6 cells	<i>Mutation Research</i>	557	1-6	2004

	発表者氏名	論文タイトル名	発表誌名	巻号	ページ	出版年
13)	Takayoshi Suzuki, Yuki Nagano, Azusa Matsuura, Arihiro Kohara, Shin-ichi Ninomiya, Kohfuku Kohda, <u>Naoki Miyata</u>	Novel histone deacetylase inhibitors: design, synthesis, enzyme inhibition, and binding mode study of SAHA-Based non-hydroxamates	<i>Bioorganic & Medicinal Chemistry Letters</i>	13(24)	4321-4326	2003
14)	Ken-ichi Saeki, Tomoaki Matsuda, Taka-aki Kato, Katsuya Yamada, Takaharu Mizutani, Saburo Matsui, Kiyoshi Fukuhara, <u>Naoki Miyata</u>	Activation of the Human Ah Receptor by Aza-Polycyclic Aromatic Hydrocarbons and Their Halogenated Derivatives	<i>Biol. Pharm. Bull.</i>	26(4)	448-452	2003
15)	Hiroshi Tokiwa, Nobuyuki Sera, Kiyoshi Fukuhara, Hideo Utsumi, Shigeki Sasaki, <u>Naoki Miyata</u>	Structural activity relationship between Salmonella-mutagenicity and nitro-orientation of nitroazaphenanthrenes	<i>Chemico-Biological Interactions</i>	146(1)	19-25	2003
16)	Yamada K., <u>Suzuki T.</u> , Hakura A., Mizutani T., Saeki K.,	Metabolic activation of 10-aza-substituted benzo[a]pyrene by cytochrome P4501A2 in human liver microsome	<i>Mutation Research</i>	557	159-165	2003
17)	Thybaud V., Dean S., Nohmi T., de Boer J., Douglas G. R., Glickman B. W., Gorelick N. J., Heddle J. A., Heflick R. H., Lambert I., Martus H. J., Mirsalis J. C., <u>Suzuki T.</u> , Yajima N.	In vivo transgenic mutation	<i>Mutation Research</i>	540	141-151	2003
18)	Itoh T., Kuwahara T., Suzuki T., Hayashi M., Ohnishi Y.	Regional mutagenicity of Heterocyclic amines in the intestine: mutation analysis of the <i>cII</i> gene in lambda/lacZ transgenic mice	<i>Mutation Research</i>	539	99-108	2003
19)	Ikuo Nakanishi, Shunichi Fukuzumi, Toshifumi Konishi, kei Ohkubo, Mamoru Fujitsuka, Osamu Ito and <u>Naoki Miyata</u>	DNA Cleavage via Superoxide Anion Formed in Photoinduced Electron Transfer from NADH to gamma-Cyclodextrin-Bicapped C60 in an Oxygen-Saturated Aqueous Solituion	<i>J. Phys. Chem. B</i>	106(9)	2372-2380	2002
20)	Kiyoshi Fukuhara, Ikuo Nakanishi, Hisao Kansui, Etsuko Sugiyama, Mitsuhiro Kimura, Tomokazu Shimada, Shiro Urano, Kentaro Yamaguchi, and <u>Naoki Miyata</u>	Enhanced Radical-Scavenging Activity of a Planar Catechin Analogue	<i>J. Am. Chem. Soc.</i>	124(21)	5952-5953	2002
21)	Matsuoka, Atsuko; Takeshita, Kenji; Furuta, Ayumi; Ozaki, Masayasu; Fukuhara, Kiyoshi; <u>Miyata, Naoki</u>	The 4'-hydroxy group is responsible for the in vitro cytogenetic activity of resveratrol	<i>Mutation Research</i>	521(1-2)	29-35	2002
22)	Yamada K, <u>Suzuki T.</u> , Kohara A, Hayashi M, Hakura A, Mizutani T, Saeki K.	Effect of 10-aza-substitution on benzo[a]pyrene mutagenicity in vivo and in vitro.	<i>Mutation Research</i>	521	187-200	2002
23)	Kohara A, <u>Suzuki T.</u> , Honma M, Ohwada T, Hayashi M	Dinitropyrenes induce gene mutations in multiple organs of the lambda/lacZ transgenic mouse (Muta Mouse)	<i>Mutation Research</i>	515	73-83	2002
24)	Kohara A, <u>Suzuki T.</u> , Honma M, Ohwada T, Hayashi M.	Mutagenicity of aristolochic acid in the lambda/lacZ transgenic mouse (MutaMouse)	<i>Mutation Research</i>	515	63-72	2002

III. 研究成果の刊行物・別刷

平成14年度～平成16年度

Novel Inhibitors of Human Histone Deacetylases: Design, Synthesis, Enzyme Inhibition, and Cancer Cell Growth Inhibition of SAHA-Based Non-hydroxamates

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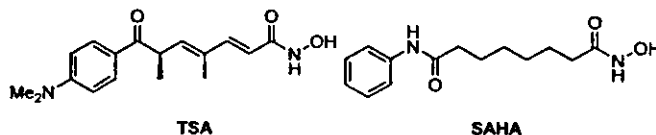
To find novel non-hydroxamate histone deacetylase (HDAC) inhibitors, a series of compounds modeled after suberoylanilide hydroxamic acid (SAHA) was designed and synthesized. In this series, compound **7**, in which the hydroxamic acid of SAHA is replaced by a thiol, was found to be as potent as SAHA, and optimization of this series led to the identification of HDAC inhibitors more potent than SAHA. In cancer cell growth inhibition assay, *S*-isobutyryl derivative **51** showed strong activity, and its potency was comparable to that of SAHA. The cancer cell growth inhibitory activity was verified to be the result of histone hyperacetylation and subsequent induction of p21^{WAF1/CIP1} by Western blot analysis. Kinetic enzyme assay and molecular modeling suggest the thiol formed by enzymatic hydrolysis within the cell interacts with the zinc ion in the active site of HDACs.

Introduction

The reversible acetylation of the ϵ -amino groups of specific histone lysine residues by histone deacetylases (HDACs) and histone acetyl transferases is an important regulatory mechanism of gene expression.¹ When HDACs are inhibited, histone hyperacetylation occurs. The disruption of the chromatin structure by histone hyperacetylation leads to the transcriptional activation of a number of genes.² One important outcome of the activation is induction of the cyclin-dependent kinase inhibitory protein p21^{WAF1/CIP1}, which causes cell cycle arrest.³ Indeed, HDAC inhibitors such as trichostatin A (TSA) and suberoylanilide hydroxamic acid (SAHA) (Chart 1) have been reported to inhibit cell growth, induce terminal differentiation in tumor cells,⁴ and prevent the formation of malignant tumors in mice.⁵ Therefore, HDACs have emerged as attractive targets in anticancer drug development, and HDAC inhibitors have also been viewed as useful tools to study the function of these enzymes.

Many groups have ongoing research programs to find nonpeptide small-molecule inhibitors of HDACs, and these efforts have led to the identification of several classes of inhibitors.⁶ Most previously reported HDAC inhibitors belong to hydroxamic acid derivatives, typified by TSA and SAHA, which are thought to chelate the zinc ion in the active site in a bidentate fashion through its CO and OH groups.⁷ However, hydroxamic acids occasionally have been associated with problems such as poor pharmacokinetics and severe toxicity.⁸ Thus, it has become increasingly desirable to find

Chart 1



replacements that possess strong inhibitory action against HDACs. In addition, in terms of biological research, the discovery of novel zinc-binding groups (ZBGs) may lead to a new type of HDAC isozyme-selective inhibitors which are useful as tools for probing the biology of the enzyme.⁹ Thus far, *o*-aminoanilide,^{9,10} electrophilic ketones,¹¹ and *N*-formyl hydroxylamine¹² have been reported as ZBGs in small-molecule HDAC inhibitors. However, most of them have reduced potency as compared to hydroxamic acid, and unfortunately, HDAC inhibitors bearing electrophilic ketones¹¹ have a metabolic disadvantage in that they are readily reduced to inactive alcohols *in vivo*, even within cells. We therefore initiated a search for replacement groups for hydroxamic acid with the goal of drug discovery as well as finding new tools for biological research, and found some potent non-hydroxamate small-molecule HDAC inhibitors.¹³ We now present a full account of our study reporting the design, synthesis, HDAC inhibition, cancer cell growth inhibition, and binding mode analysis of non-hydroxamates based on the structure of SAHA.

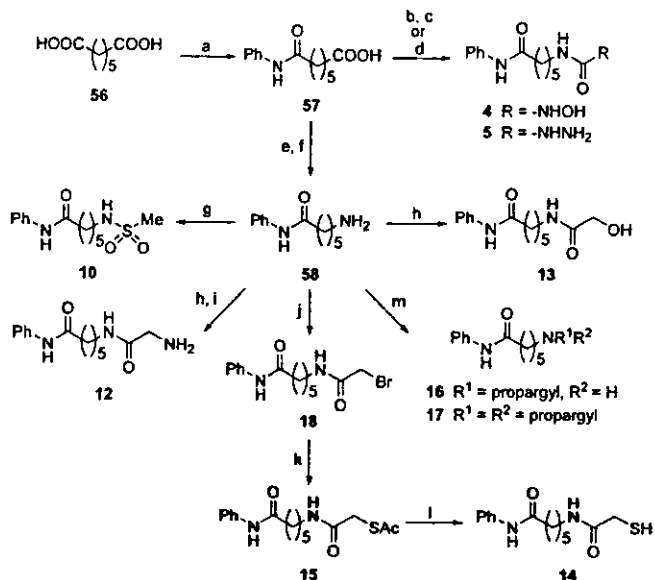
Chemistry

The compounds prepared for this study are shown in Tables 1–5. The routes used for synthesis of the compounds are shown in Schemes 1–4. Scheme 1 shows the preparation of compounds **4**, **5**, **10**, **12–17**, and **18**. Compounds **4** and **5** were synthesized from pimelic acid **56**. The condensation of pimelic acid **56** with an equiva-

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[†] Graduate School of Pharmaceutical Sciences, Nagoya City University.

[‡] Nippon Kayaku Co., Ltd.

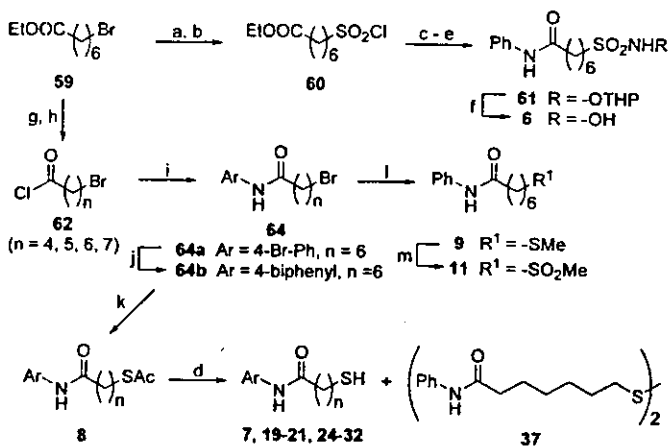
Scheme 1^a

^a Reagents and conditions: (a) aniline, 180 °C; (b) diphenylphosphoryl azide (DPPA), Et₃N, toluene, reflux, and then *O*-(2-tetrahydropyran-2-yl)hydroxylamine, reflux; (c) TsOH, MeOH, rt; (d) DPPA, Et₃N, benzene, reflux, and then hydrazine monohydrate, reflux; (e) DPPA, Et₃N, benzene, reflux, and then BnOH, reflux; (f) H₂, 5% Pd-C, MeOH, rt; (g) MsCl, pyridine, rt; (h) HOCH₂COOH or BocNHCH₂COOH, EDCI, HOBt, DMF, rt; (i) TFA, CHCl₃, rt; (j) BrCH₂COBr, Et₃N, THF, rt; (k) AcSK, EtOH, rt; (l) K₂CO₃, MeOH, rt; (m) propargyl bromide, K₂CO₃, MeOH, rt.

lent amount of aniline gave mono-anilide 57. Curtius rearrangement of the acyl azide prepared from carboxylic acid 57 using diphenylphosphoryl azide provided the isocyanates, which on treatment with *O*-tetrahydropyran-2-yl (THP) hydroxylamine or hydrazine gave *O*-THP hydroxyurea and semicarbazide 5. Deprotection of the THP group of the *O*-THP hydroxyurea under acidic conditions afforded hydroxyurea 4.

Compounds 10, 12–17, and 18 were prepared from carboxylic acid 57 obtained above via amine 58 by the procedure outlined in Scheme 1. Carboxylic acid 57 was converted to amine 58 with a three-step sequence: Curtius rearrangement of the acyl azide prepared from carboxylic acid 57, treatment of the resulting isocyanates with benzyl alcohol, and removal of the Z group by hydrogenation. Coupling between amine 58 and methanesulfonyl chloride afforded sulfonamide 10. The reaction of amine 58 with *N*-Boc glycine in the presence of EDCI and HOBt in DMF was followed by treatment with trifluoroacetic acid to give aminoacetamide 12. Hydroxyacetamide 13 was obtained in one step using the procedure described for 12. The amino group of 58 was acylated with bromoacetyl bromide to yield bromoacetamide 18. Bromide 18 was treated with potassium thioacetate to give thioacetate 15, after which deacetylation of the thioacetate in the presence of K₂CO₃ in MeOH gave mercaptoacetamide 14. The amine 58 was allowed to react with propargyl bromide in the presence of K₂CO₃ to give mono- and di-alkylated compounds 16 and 17.

Compounds 6–9, 11, 19–21, 24–32, and 37 were prepared from another starting material, 59 (Scheme 2). The preparation of hydroxysulfonamide 6 was achieved via sulfonyl chloride 60. Bromide 59 was converted to sulfonyl chloride 60 by sulfation and by a

Scheme 2^a

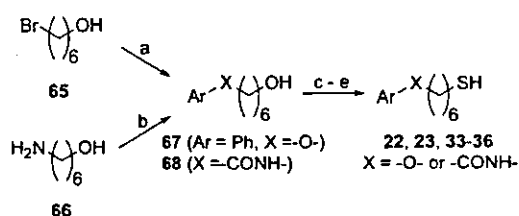
^a Reagents and conditions: (a) Na₂SO₃, EtOH, H₂O, reflux; (b) SOCl₂, DMF, toluene, reflux; (c) *O*-(2-tetrahydropyran-2-yl)hydroxylamine, 4-(dimethylamino)pyridine, pyridine, CH₂Cl₂, rt; (d) 2N aq NaOH, EtOH, rt; (e) aniline, EDCI, HOBt, DMF, rt; (f) TFA, CH₂Cl₂, 60 °C; (g) LiOH·H₂O, EtOH, THF, H₂O, rt; (h) (COCl)₂, DMF, CH₂Cl₂, rt; (i) ArNH₂ (63), Et₃N, CH₂Cl₂, rt; (j) PhB(OH)₂, Pd(PPh₃)₄, NaHCO₃, 1-methyl-2-pyrrolidinone, H₂O, 80 °C; (k) AcSK, EtOH, rt; (l) 15% aq NaSMe, EtOH, rt; (m) *m*-chloroperoxybenzoic acid, CH₂Cl₂, rt.

subsequent reaction with thionyl chloride. The sulfonyl chloride 60 was treated with *O*-THP hydroxylamine to give *O*-THP hydroxysulfonamide, after which hydrolysis of the ester under alkaline conditions and subsequent amide formation with aniline gave compound 61. Removal of the THP group of compound 61 by treatment with trifluoroacetic acid gave hydroxysulfonamide 6.

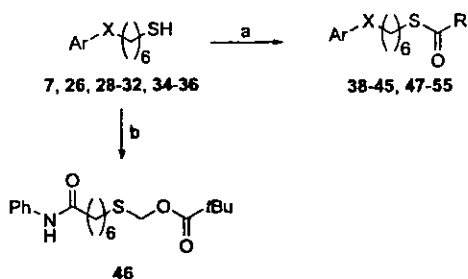
Compounds 7–9, 11, 19–21, 24–32, and 37 were synthesized from the corresponding acid chlorides 62 (62a (n = 4) and 62b (n = 5) are commercially available) by the route shown in Scheme 2. 62c (n = 6) was prepared from ester 59 by hydrolysis of the ethyl ester and a subsequent reaction with oxalyl chloride, and 62d (n = 7) was obtained in the same way as 62c. The amino group of aromatic amines 63 was acylated with an appropriate acid chloride 62 to give the amides 64. Suzuki coupling¹⁴ of bromobenzene 64a with phenylboronic acid provided the biphenyl 64b. Bromides 64 were treated with potassium thioacetate to give compound 8, after which hydrolysis of the thioacetates under alkaline conditions gave the desired compounds 7, 19–21, 24–31, and 32, and disulfide 37 was obtained as a byproduct when thiol 7 was synthesized. Sulfide 9 was prepared by the alkylation of methylmercaptan with bromide 64c (Ar = Ph, n = 6). Oxidation of 9 with 2 equiv of *m*-chloroperoxybenzoic acid provided the sulfone 11.

Thiols 22, 23, 33–35, and 36 were prepared from alcohol 65 or 66 by the procedure outlined in Scheme 3. Treatment of bromide 65 with phenol in the presence of K₂CO₃ gave ether 67, and condensation of amine 66 with an appropriate aromatic carboxylic acid 69 afforded amides 68. Alcohols 67 and 68 were converted to thiols 22, 23, 33–35, and 36 in three steps by conversion of the alcohols to bromides, treatment of the bromides with potassium thioacetate, and hydrolysis of the resulting thioacetates.

The preparation of *S*-chemically modified compounds 38–54, and 55 is shown in Scheme 4. Thiols 7, 26, 28–32, 34, 35, and 36 were coupled with the corresponding

Scheme 3^a

^a Reagents and conditions: (a) Phenol, K₂CO₃, DMF, 80 °C; (b) ArCOOH (69), EDCI, HOBT, DMF, rt; (c) CBr₄, PPh₃, CH₂Cl₂, 0 °C; (d) AcSK, EtOH, rt; (e) 2N aq NaOH, EtOH, THF, rt.

Scheme 4^a

^a Reagents and conditions: (a) RCOCl (70), 4-(dimethylamino)pyridine, pyridine, CH₂Cl₂, rt; (b) NaH, chloromethyl pivalate, DMF, 0 °C to room temperature.

acyl chloride 70 to give thioesters 38–45, 47–54, and 55. Alkylation of thiol 7 with chloromethyl pivalate in the presence of sodium hydride in DMF afforded compound 46.

Results and Discussion

Enzyme Assays. The compounds synthesized in this study were tested with an *in vitro* assay using a HeLa nuclear extract rich in HDAC activity. The results are summarized in Tables 1–3.

The IC₅₀ values of SAHA and *o*-aminoanilide 1 were 0.28 μM and 120 μM, respectively (entries 1 and 2). α -Ketoamide 2 and *N*-formyl hydroxylamine 3 were reported previously to inhibit HDACs with an IC₅₀ of 0.34 μM and 2.8 μM, respectively (entries 3 and 4).^{11b,12}

The crystal structures of an archaeobacterial HDAC homologue (HDAC-like protein, HDLP)/hydroxamates and HDAC8/hydroxamates complexes made it clear that the hydroxamic acid group coordinates the zinc ion in the active site through its CO and OH groups and also forms three hydrogen bonds between its CO, NH, and OH groups and Tyr 306, His 143, and His 142 (HDAC8 numbering), respectively.⁷ From these data, hydroxyurea 4, semicarbazide 5, and hydroxysulfonamide 6 were synthesized and tested as HDAC inhibitors because it is possible for them to chelate zinc ion and form hydrogen bonds with Tyr and His like SAHA. Among these three compounds, hydroxyurea 4 and semicarbazide 5 showed anti-HDAC activity and the IC₅₀ values were comparable to that of *o*-aminoanilide 1 (entries 5, 6, and 7). However, they were much less effective than SAHA.

Thiols seemed to be reasonable targets for hydroxamic acid replacements, because zinc ion is highly thiophilic and thiol derivatives have been reported to inhibit zinc-dependent enzymes such as angiotensin converting enzyme¹⁵ and matrix metalloproteinases.¹⁶ Furthermore, macrocyclic compounds bearing a disulfide group such as FK228 have been reported recently to inhibit

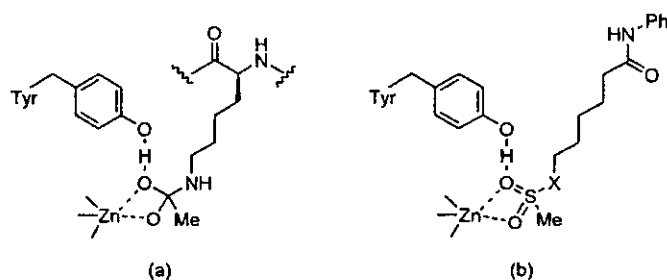


Figure 1. The transition state proposed for HDACs (a), and models for the binding of sulfone derivatives (b).

HDACs under reductive conditions.¹⁷ Surprisingly, although the inhibitory ability of monodentate ZBGs such as thiol was thought to be less than that of bidentate ZBGs such as hydroxamate, *N*-formyl hydroxylamine, and hydrated electrophilic ketones,¹⁶ the activity of thiol 7 was far greater than expected. A pronounced inhibitory effect (IC₅₀ = 0.21 μM) was observed with thiol 7, which was much more active than previously reported non-hydroxamates such as *o*-aminoanilide, *N*-formyl hydroxylamine, and trifluoromethyl ketone,^{11a} and as potent as α -ketoamide 2 and SAHA (entry 8). To confirm that the thiol group plays an important role in anti-HDAC activity, thioacetate 8a and methyl sulfide 9 were tested. As expected, thiol transformation into thioacetate and methyl sulfide led to an inhibitor that was about 30-fold less potent and a compound devoid of anti-HDAC activity, respectively (entries 9 and 10). These results suggest that thiolate anion generated under physiological conditions has an intimate involvement in the interaction with the zinc ion in the active site.

The crystal structures of the HDLP/hydroxamates and HDAC8/hydroxamates complexes have led to a solid understanding of not only the three-dimensional structure of the active site of HDACs but also the catalytic mechanism for the deacetylation of acetylated lysine substrate.⁷ It has been proposed that the carbonyl oxygen of this substrate could bind the zinc, and the carbonyl could be attacked by a zinc-chelating water molecule (Figure 2a), which would result in the production of deacetylated lysine via a tetrahedral carbon-containing transition state (Figure 1a). On the basis of the proposed catalytic mechanism, we attempted to design non-hydroxamate HDAC inhibitors. First, we designed transition-state (TS) analogues. The TS of HDAC deacetylation was estimated to include a tetrahedral carbon (Figure 1a) as with other zinc proteases.¹⁸ We focused attention on sulfone derivative TS analogues because it has been suggested that the sulfonamide moiety has strong similarity with the TS of amide bond hydrolysis, both from a steric and an electronic point of view.¹⁹ Compounds 10 and 11, in which a hydroxamic acid of SAHA is replaced by a sulfonamide and a sulfone, respectively, were designed and synthesized as TS analogues (Figure 1b). Of these two TS analogues, sulfone 11 showed anti-HDAC activity and the IC₅₀ value was 230 μM (entries 11 and 12). However, sulfone 11 was approximately 820-fold less effective than SAHA.

Our next approach was based on the proposed deacetylation mechanism whereby a zinc-chelating water molecule activated by His142 and His 143 (HDAC8 numbering) makes a nucleophilic attack on the carbonyl carbon of an acetylated lysine substrate (Figure 2a).

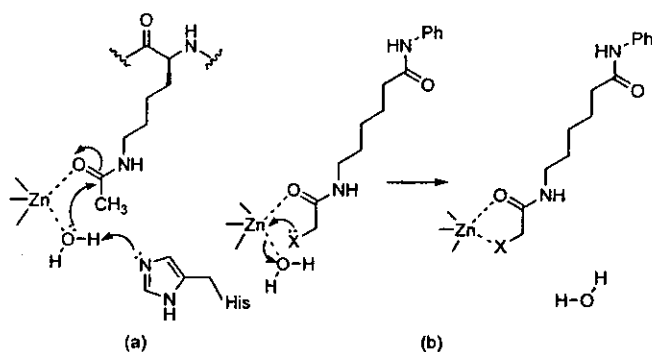


Figure 2. The mechanism proposed for the deacetylation of acetylated lysine substrate (a), and a model for the binding of heteroatom-containing substrate analogues to zinc ion (b).

Table 1. HDAC Inhibition Data for SAHA and SAHA-based Non-hydroxamates^a

entry	compd	R	n	% inhibn at 100 μM	IC ₅₀ (μM)
1	SAHA ^b	-CONHOH	6	100	0.28
2	1 ^c		6	48	120
3	2	-COCONHMe	6	ND	0.34 ^d
4	3		7	ND	2.8 ^e
5	4	-NHCONHOH	5	58	80
6	5	-NHCONHNH ₂	5	35	150
7	6	-SO ₂ NHOH	6	14	>100
8	7	-SH	6	100	0.21
9	8a	-SAc	6	85	7.1
10	9	-SMc	6	11	>100
11	10	-NHSO ₂ Me	5	10	7500
12	11	-SO ₂ Me	6	33	230
13	12 ^f	-NHCOCH ₂ NH ₂	5	6	>100
14	13	-NHCOCH ₂ OH	5	0	>100
15	14	-NHCOCH ₂ SH	5	99	0.39
16	15	-NHCOCH ₂ SAc	5	72	22
17	16		5	14	>100
18	17		5	0	>100
19	18	-NHCOCH ₂ Br	5	78	17

^a Values are means of at least three experiments. ^b Prepared as described in ref 26. ^c Prepared as described in ref 9a. ^d Data taken from the literature (ref 11b). ^e Data taken from the literature (ref 12). ^f Trifluoroacetic acid salt. ND = No data.

With this mechanism, if the water molecule is forcibly removed from the zinc ion, the HDACs would supposedly be inhibited. We then designed and synthesized heteroatom-containing substrate analogues 12, 13, and 14. These analogues would be recognized as substrates by HDACs and would be easily taken into the active site where they could force the water molecule off the zinc ion and the reactive site for the deacetylation by chelation of the heteroatom to the zinc ion, and might behave as HDAC inhibitors (Figure 2b). As shown in Table 1 (entries 13, 14, and 15), potent inhibition was observed with mercaptoacetamide 14, while 12 and 13 did not possess HDAC inhibitory activities. Mercaptoacetamide 14 exhibited an IC₅₀ of 0.39 μM, and its activity largely surpassed those of *o*-aminoanilide 1 and *N*-formyl hydroxylamine 3 and was comparable to those of α -ketoamide 2 and SAHA. As expected, thiol trans-

Table 2. Effect of Linker Variation on HDAC Inhibitory Activity of Thiols^a

entry	compd	X	n	IC ₅₀ (μM)
1	7	-NHCO-	6	0.21
2	19	-NHCO-	7	1.5
3	20	-NHCO-	5	0.37
4	21	-NHCO-	4	6.2
5	22	-O-	6	11
6	23	-CONH-	6	0.36

^a Values are means of at least three experiments.

Table 3. Effect of Aromatic Group Variation on HDAC Inhibitory Activity of Thiols^a

entry	compd	Ar	X	IC ₅₀ (μM)
1	7	-Ph	-NHCO-	0.21
2	24		-NHCO-	1.2
3	25		-NHCO-	1.1
4	26		-NHCO-	0.075
5	27		-NHCO-	0.62
6	28		-NHCO-	0.21
7	29		-NHCO-	0.11
8	30		-NHCO-	0.072
9	31		-NHCO-	0.17
10	32		-NHCO-	0.34
11	23	-Ph	-CONH-	0.36
12	33		-CONH-	0.61
13	34		-CONH-	0.085
14	35		-CONH-	0.079
15	36		-CONH-	0.10

^a Values are means of at least three experiments.

formation into thioacetate (15) led to a 55-fold less potent inhibitor. This result suggests the ease of ionization of thiol is an important factor for HDAC inhibition like the case of thiol 7.

We turned our attention to irreversible HDAC inhibitors. TPX B is an irreversible HDAC inhibitor,²⁰ and finding more specific and simpler irreversible HDAC inhibitors is useful for the isolation and cloning of an HDAC.² As described above, the crystal structures of the HDLP/hydroxamates and HDAC8/hydroxamates complexes revealed that the hydroxamic acid group forms three hydrogen bonds with Tyr 306, His 143, and His 142, and furthermore, zinc ion is coordinated by His 180, Asp 178, and Asp 267 (HDAC8 numbering). Since the phenol group of Tyr, the imidazole group of His, and the carboxyl group of Asp are able to react with electrophiles, we prepared analogues bearing propargyl

Table 4. Cell Growth Inhibition Data on NCI-H460 Cells for Compound **7** and Its S-Modified Prodrugs^a

entry	compd	R	EC ₅₀ (μM)
1	7	-H	>50 ^b
2	37		>50 ^c
3	8a	-Ac	36
4	38	-COEt	28
5	39	-CO <i>n</i> -Pr	22
6	40	-CO <i>i</i> -Pr	20
7	41	-CO <i>t</i> -Bu	>50 ^d
8	42		27
9	43		21
10	44	-Bz	25
11	45		24
12	46	-CH ₂ OCO <i>t</i> -Bu	25

^a Values are means of at least two experiments. ^b 34% inhibition at 50 μM. ^c 10% inhibition at 50 μM. ^d 42% inhibition at 50 μM.

amino (**16**, **17**) and bromoacetamide (**18**) which could form covalent bonds with Tyr, His, and Asp of the enzyme, and evaluated their anti-HDAC activities. While propargyl amino compounds **16** and **17** did not possess HDAC inhibitory activities, more potent inhibition was observed with bromoacetamide **18** (entries 17, 18, and 19). Bromoacetamide **18** exhibited an IC₅₀ of 17 μM and its activity was about 9-fold as strong as that of *o*-aminoanilide **1**, but much weaker than that of SAHA.

With the results shown in Table 1, we were encouraged to study further the structure-activity relationship (SAR) and structural optimization. We selected thiol **7** for further study.²¹ First, we examined the effect of linker parts of thiol **7**. The results are summarized in Table 2. HDAC inhibition was distinctly dependent on chain length, with *n* = 7 (**19**) and *n* = 4 (**21**) resulting in less potent inhibitors. However, compound **20**, in which *n* = 5, showed essentially the same potency as compound **7**, in which *n* = 6 (entries 1–4). The similar SAR between thiols and hydroxamates, with *n* = 6 optimal,²² indicates that thiols inhibit HDACs in a binding mode similar to that of hydroxamates. As for the group attaching the phenyl moiety, ether **22** displayed moderate activity, whereas the activity of the reversed amide **23** was maintained (entries 5 and 6).

Next, the aromatic group was examined (Table 3). In the amide-linked series (entries 1–10), 4-substituted phenyl compounds tended to decrease the potency. Specifically, compounds **24** (Ar = 4-NMe₂-Ph), **25** (Ar = 4-biphenyl), and **27** (Ar = 4-PhO-Ph) showed about a 3- to 6-fold decrease in potency when compared to the parent thiol **7** (entries 2, 3, and 5). On the other hand, compound **26**, in which a phenyl group was introduced at the 3-position of the phenyl group of **7**, showed 3-fold increased inhibitory activity (IC₅₀ of 0.075 μM, entry 4). In addition, 3-phenoxy compound **28** was equipotent with compound **7** (entry 6). We investigated the effect of the replacement of the phenyl group of compound **7** with heteroaryl rings (entries 7, 8, 9, and 10). Changing the benzene ring to a 3-pyridine ring (**29**), 4-phenyl-2-thiazole ring (**31**), and 2-benzothiazole ring (**32**) sus-

Table 5. Cell Growth Inhibition Data on NCI-H460 Cells for Compound **40** and Its Derivatives^a

entry	compd	Ar	X	EC ₅₀ (μM)
1	40	-Ph	-NHCO-	20
2	47		-NHCO-	2.8
3	48		-NHCO-	25
4	49		-NHCO-	2.9
5	50		-NHCO-	8.0
6	51		-NHCO-	2.1
7	52		-NHCO-	9.5
8	53		-CONH-	12
9	54		-CONH-	4.1
10	55		-CONH-	12

^a Values are means of at least two experiments.

Table 6. Growth Inhibition of Various Cancer Cells Using SAHA and Compound **51**^a

cell	SAHA, EC ₅₀ (μM)	51 , EC ₅₀ (μM)	
MDA-MB-231	breast cancer	1.5	2.3
SNB-78	central nervous system	16	9.1
HCT116	colon cancer	0.58	3.0
NCI-H226	lung cancer	2.6	2.6
LOX-IMVI	melanoma	1.3	1.1
SK-OV-3	ovarian cancer	2.5	4.5
RXF-631L	renal cancer	2.0	2.4
St-4	stomach cancer	5.2	5.0
DU-145	prostate cancer	1.6	4.5
mean		3.7	3.8

^a Values are means of at least two experiments.

tained or slightly reduced the activity, whereas quinoline **30** had improved activity (IC₅₀ of 0.072 μM), and turned out to be the most potent compound in this series. The reverse amide-linked series (entries 11–15) exhibited potencies similar to or greater than the parent thiol **23**, with the exception of **33** (Ar = 4-NMe₂-Ph), which resulted in a slightly less potent inhibitor. In particular, the reversed amides **34** with a naphthalene substituent and **35** with a benzofuran substituent exhibited about 3-fold increases in potency (IC₅₀s of 0.085 μM and 0.079 μM, respectively). As a result, IC₅₀s in the double-digit nanomolar range were observed with 3-biphenyl **26**, quinoline **30**, naphthalene **34**, and benzofuran **35**, which were approximately 3- to 4-fold more potent than SAHA.

Cancer Cell Growth Inhibition Assay. To confirm the effectiveness of thiol-based HDAC inhibitors as anticancer drugs and tools for biological research, thiol **7** was initially tested in a cancer cell growth inhibition

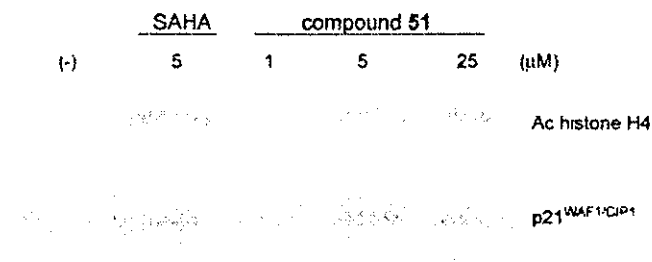


Figure 3. Western blot analysis of histone hyperacetylation and p21^{WAF1/CIP1} induction in HCT 116 cells produced by compound 51 and by reference compound SAHA.

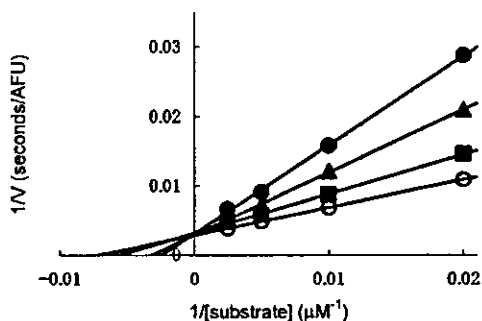


Figure 4. Reciprocal rate vs reciprocal acetylated lysine substrate concentration in the presence of 0.3 (●), 0.1 (▲), 0.03 (■), and 0 (○) μM of 7.

assay using human lung cancer NCI-H460 cells against which it was found to be only weakly potent, although 7 was highly active in an enzymatic assay (Table 4, entry 1). The reason for the weak activity of thiol 7 is unclear, but it is reasonable to assume that it was due to poor membrane permeability resulting from the highly polar character of this compound, and that a transient masking of the sulfhydryl group could improve its permeability and its ability to inhibit cancer cell growth. Therefore, we investigated the possibility of improving the inhibition using the prodrug approach. In the search for a suitable prodrug of thiols, disulfides

seemed to be attractive targets, because it has been reported that the disulfide bond of macrocyclic compounds bearing a disulfide group such as FK228 is reduced in the cellular environment, releasing the free thiol analogue as the active species.¹⁷ However, contrary to our expectation, disulfide 37 failed to exhibit a growth inhibitory effect on NCI-H460 cells (entry 2). Next, we examined the activity of acetyl compound 8a. Acetyl compound 8a proved to be relatively potent compared with thiol 7 and disulfide 37 (EC_{50} of 36 μM) (entry 3). This result suggests that 8a permeates the cell membrane more efficiently than thiol 7, and is converted to thiol 7 by enzymatic hydrolysis within the cell.²³ Encouraged by this finding, we prepared other *S*-acyl compounds (38–45) and evaluated their activities (entries 4–11). This series of compounds exhibited greater potency than acetyl compound 8a, except for pivaloyl compound 41, which was a less potent inhibitor. In particular, isobutyryl compound 40 showed about a 2-fold increase in activity when compared to acetyl compound 8a (EC_{50} of 20 μM). The compound bearing a (pivaloyloxy)methyl group²⁴ (46) was slightly less active than isobutyryl compound 40 (entry 12).

With the results shown in Table 4, a selected set of active compounds from the enzymatic assay was *S*-isobutyrylated and tested as cancer cell growth inhibitors (Table 5). Much to our satisfaction, changing the phenyl group of compound 40 to other aromatic groups led to positive results. Isobutyryl analogues 47–55 were generally more potent than the parent compound 40; the sole exception is 48 (Ar = 3-OPh-Ph) which was slightly less active than compound 40 (entry 3). Notably, 3-biphenyl (47), 3-pyridinyl (49), and 4-phenyl-2-thiazolyl (51) analogues showed strong activity in inhibiting the growth of NCI-H460 cells, with EC_{50} s of 2–3 μM . Furthermore, we evaluated cancer cell growth inhibition by SAHA and 51, the most potent compound in this study, against nine other human cancer cell lines

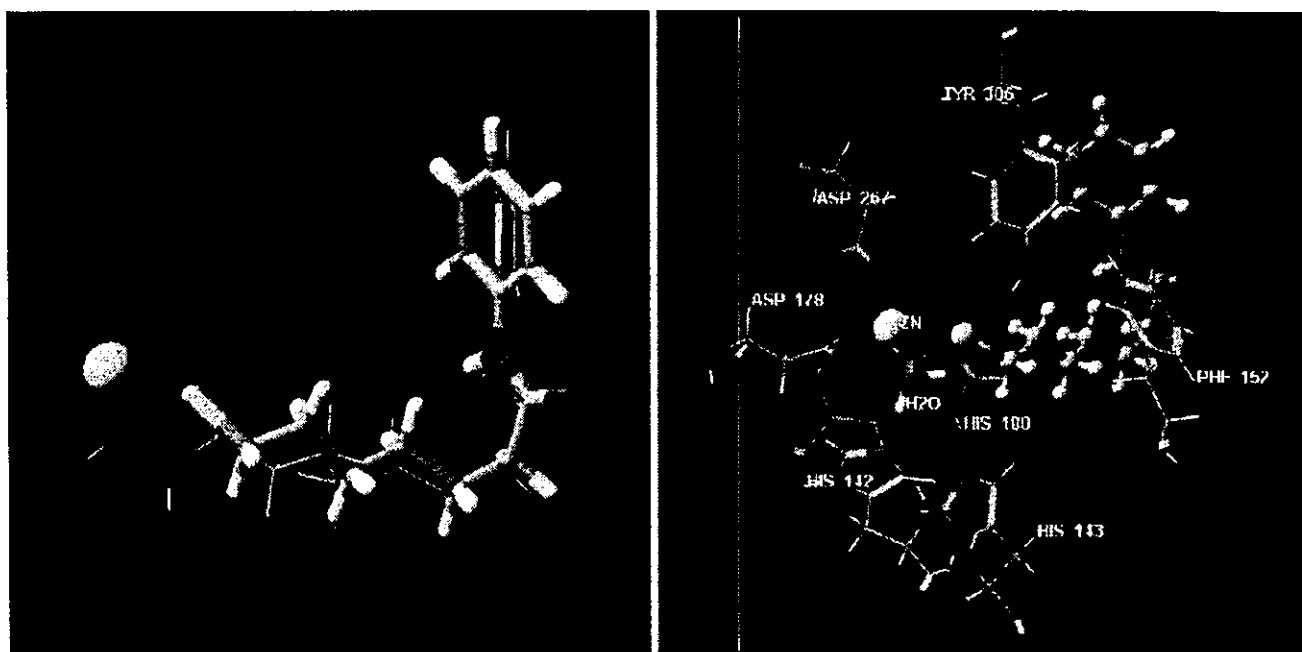


Figure 5. Superposition of the low energy conformations of 7 (tube) and SAHA (wire) (left). The HDAC8 pocket is not shown for the sake of clarity. View of the conformation of 7 (ball-and-stick) docked in the HDAC8 catalytic core (right). Residues around the zinc ion and a water molecule are displayed as wires and tubes, respectively.

(Table 6). Compound **51** exerted potent growth inhibition against various human cancer cells, with EC_{50} values ranging from 1 to 10 μM , and these inhibitory activities were comparable to those of SAHA (average EC_{50} of **51** 3.8 μM , SAHA 3.7 μM) which is currently being evaluated in clinical trials for use in the treatment of cancer.

By Western blot analysis, cancer cell growth inhibition with compound **51** was verified to be the result of inhibition of HDACs (Figure 3). Treatment of HCT 116 cells with compound **51** gave rise to elevated and dose-dependent levels of acetylated histone H4 and p21^{WAF1/CIP1}.

Inhibitory Mechanism Study. Since the results of cancer cell growth inhibition and Western blot analysis have suggested that thiols generated from *S*-acyl prodrugs by enzymatic hydrolysis within the cell inhibit intracellular HDACs, we next studied the mechanism by which thiols inhibit HDACs in greater detail. Although the sulfhydryl group of thiol derivatives was designed as a ZBG, it is possible that thiols inhibit HDACs by forming a covalent disulfide bond with cysteine residues on these enzymes. We examined this possibility using a double reciprocal plot of $1/V$ versus $1/[\text{substrate}]$ at varying concentrations of inhibitor **7** (Figure 4), and the data from this study established that thiol **7** engages in competitive inhibition versus acetylated lysine substrate, with an inhibition constant (K_i) of 0.11 μM . Since cysteine is not a component in the construction of the active site of HDACs, the sulfhydryl group of **7** likely interacts with the zinc in the active site.

Binding Mode Study. Since thiol **7** proved to be a competitive inhibitor and to act within the active center of HDACs, we studied its binding mode within this site. The low energy conformations of **7** and SAHA were calculated when docked in the model based on the crystal structure of HDAC8 (PDB code 1T64, 1T67, 1T69, and 1VKG) using Macromodel 8.1 software.²⁵ The anilide group and alkyl chain of **7** and SAHA were essentially superimposed in the binding pocket, and the binding mode of **7** was found to be similar to that of SAHA (Figure 5, left). An inspection of the HDAC8/**7** complex shows that the sulfur atom of **7** was located 2.35 Å from the zinc ion, 2.24 Å from the OH group of Tyr 306, and 2.66 Å from a water molecule which forms a hydrogen bond with the imidazole group of His142 (Figure 5, right). This suggests that thiols strongly inhibit HDACs by interacting directly with zinc ion, Tyr 306, and His 142 via a water molecule.

Conclusion

We have designed and prepared a series of SAHA-based compounds as (i) hydroxamic acid mimics by structure-based drug design (compounds 4–6), (ii) thiol-based analogues (compounds 7–9), (iii) transition-state analogues (compounds 10 and 11), (iv) heteroatom-containing substrate analogues by mechanism-based drug design (compounds 12–15), and (v) irreversible inhibition-oriented compounds (compounds 16–18), and evaluated their inhibitory effect on HDACs. In this series, thiol **7** and mercaptoacetamide **14** were found to be much more potent HDAC inhibitors than previously reported non-hydroxamates, and as potent as

α -ketoamide **2** and SAHA. At present, thiol is one of the most active ZBG among small-molecule HDAC inhibitors. Optimization of thiol derivatives led to the identification of inhibitors more effective than SAHA (compounds **26**, **30**, **34**, and **35**). We have also identified a potent cancer cell growth inhibitor, compound **51**, by the prodrug formation of thiol-based HDAC inhibitors. Thiol **7** exhibits strong competitive inhibition of an acetylated lysine substrate, and molecular modeling suggests that the thiol interacts with zinc, Tyr 306, and His 142 (HDAC8 numbering) in the active site.

In conclusion, we have identified several new lead structures including thiol, from which more potent HDAC inhibitors can be developed. As far as we could determine, this is the first systematic study of ZBGs for HDAC inhibitors. We believe that the findings of this study should be of value in future studies for the development of ideal anticancer drugs and tools for biological research such as HDAC isozyme-selective inhibitors.

Experimental Section

Chemistry. Melting points were determined using a Yanagimoto micro melting point apparatus or a Büchi 545 melting point apparatus and were left uncorrected. Proton nuclear magnetic resonance spectra (¹H NMR) were recorded on a JEOL JNM-LA400 or JEOL JNM-LA500 spectrometer in solvent as indicated. Chemical shifts (δ) are reported in parts per million relative to the internal standard tetramethylsilane. Elemental analysis was performed with a Yanaco CHN CORDER NT-5 analyzer, and all values were within $\pm 0.4\%$ of the calculated values. High-resolution mass spectra (HRMS) were recorded on a JEOL JMS-SX102A mass spectrometer. GC-MS analyses were performed on a Shimadzu GCMS-QP2010. Reagents and solvents were purchased from Aldrich, Tokyo Kasei Kogyo, Wako Pure Chemical Industries, and Kanto Kagaku and used without purification. Flash column chromatography was performed using silica gel 60 (particle size 0.046–0.063 mm) supplied by Merck.

6-(3-Hydroxyureido)hexanoic Acid Phenylamide (4). **Step 1: Preparation of 6-Phenylcarbamoylhexanoic Acid (57).** A mixture of aniline (5.80 g, 62.3 mmol) and pimeric acid (**56**, 10.0 g, 62.4 mmol) was stirred at 180 °C for 1 h. After cooling, the mixture was diluted with AcOEt-THF and the slurry was filtered. The filtrate was washed with saturated aqueous NaHCO₃, and the aqueous layer was acidified with concentrated HCl. The precipitated crystals were collected by filtration to give 7.11 g (49%) of **57** as a white solid: ¹H NMR (DMSO-*d*₆, 400 MHz, δ ; ppm) 11.97 (1H, broad s), 9.83 (1H, s), 7.58 (2H, d, $J = 7.8$ Hz), 7.27 (2H, t, $J = 7.9$ Hz), 7.01 (1H, t, $J = 7.4$ Hz), 2.67 (2H, t, $J = 7.4$ Hz), 2.21 (2H, t, $J = 7.3$ Hz), 1.62–1.49 (4H, m), 1.34–1.27 (2H, m).

Steps 2 and 3: Preparation of 6-(3-Hydroxyureido)hexanoic Acid Phenylamide (4). To a suspension of **57** (958 mg, 4.07 mmol) obtained above and triethylamine (744 mg, 7.35 mmol) in toluene (10 mL) was added diphenylphosphoryl azide (1.75 g, 6.34 mmol), and the mixture was heated at reflux temperature for 1 h. Next, *O*-(2-tetrahydropyranyl)hydroxylamine (380 mg, 3.11 mmol) was added, and the reaction mixture was stirred at reflux temperature for 18 h. It was then concentrated in vacuo, and the residue was dissolved in AcOEt. The AcOEt solution was washed with water, saturated aqueous NaHCO₃, and brine and was dried over Na₂SO₄. Filtration and concentration in vacuo and purification by silica gel flash chromatography (*n*-hexane/AcOEt = 1/2) gave 988 mg (69%) of the *O*-(2-tetrahydropyranyl)hydroxyurea as a white solid: ¹H NMR (CDCl₃, 500 MHz, δ ; ppm) 7.53 (2H, d, $J = 7.9$ Hz), 7.32 (2H, t, $J = 7.8$ Hz), 7.26 (1H, broad s), 7.10 (1H, t, $J = 7$ Hz), 7.05 (1H, broad s), 6.06 (1H, broad s), 4.75 (1H, d, $J = 3.6$ Hz), 3.93 (1H, m), 3.57 (1H, m), 3.33–3.26 (2H, m), 2.38 (2H, t, $J = 7.5$ Hz), 1.82–1.77 (4H, m), 1.61–1.55 (6H, m), 1.44 (2H, quintet, $J = 7.3$ Hz).

To a solution of the *O*-(2-tetrahydropyranyl)hydroxyurea (185 mg, 0.53 mmol) obtained above in MeOH (2 mL) was added 4-toluenesulfonic acid monohydrate (15 mg, 0.079 mmol). The solution was stirred overnight at room temperature, and the precipitated crystals were collected by filtration to give 46 mg (32%) of 4 as a white solid. The solid was recrystallized from MeOH-AcOEt and collected by filtration to give 34 mg of 4 as a colorless crystal: mp 148–149 °C; ¹H NMR (DMSO-*d*₆, 500 MHz, δ; ppm) 9.93 (1H, s), 8.58 (1H, s), 8.29 (1H, s), 7.65 (2H, d, *J* = 8 Hz), 7.35 (2H, t, *J* = 7.9 Hz), 7.08 (1H, t, *J* = 7.3 Hz), 6.75 (1H, t, *J* = 6 Hz), 3.10 (2H, q, *J* = 6.7 Hz), 2.36 (2H, t, *J* = 7.5 Hz), 1.65 (2H, quintet, *J* = 7.5 Hz), 1.50 (2H, quintet, *J* = 7.2 Hz), 1.34 (2H, quintet, *J* = 7.6 Hz); Anal. (C₁₃H₁₉N₃O₃) C, H, N.

6-(3-Aminoareido)hexanoic Acid Phenylamide (5). Compound 5 was prepared from 57 obtained above by using the procedure described for 4 (step 2) in 52% yield. In this case, hydrazine monohydrate was used instead of *O*-(2-tetrahydropyranyl)hydroxyurea: mp 146–147 °C; ¹H NMR (DMSO-*d*₆, 400 MHz, δ; ppm) 9.83 (1H, s), 7.58 (2H, d, *J* = 7.8 Hz), 7.27 (2H, t, *J* = 7.9 Hz), 7.01 (1H, t, *J* = 7.3 Hz), 6.83 (1H, broad s), 6.28 (1H, broad s), 4.03 (2H, broad s), 3.01 (2H, q, *J* = 6.7 Hz), 2.29 (2H, t, *J* = 7.4 Hz), 1.60–1.57 (2H, m), 1.40–1.38 (2H, m), 1.32–1.28 (2H, m); MS (EI) *m/z*: 264 (M⁺); Anal. (C₁₃H₂₀N₄O₂) C, H, N.

6-Methanesulfonylaminohexanoic Acid Phenylamide (10). Steps 1 and 2: Preparation of 6-Aminohexanoic Acid Phenylamide (58). To a suspension of 57 (1.11 g, 4.73 mmol) obtained above and triethylamine (699 mg, 6.90 mmol) in benzene (3 mL) was added diphenylphosphoryl azide (1.83 g, 6.64 mmol), and the mixture was heated at reflux temperature for 1 h. Next, benzyl alcohol (1.20 mL, 11.6 mmol) was added, and the reaction mixture was stirred at reflux temperature for 24 h. It was then concentrated in vacuo and the residue was dissolved in AcOEt. The AcOEt solution was washed with 0.4 N aqueous HCl, water, saturated aqueous NaHCO₃, and brine and was dried over Na₂SO₄. Filtration and concentration in vacuo and purification by recrystallization from CHCl₃-*n*-hexane gave 1.01 g (63%) of (6-phenylcarbamoylpentyl)carbamic acid benzyl ester as a colorless needle: ¹H NMR (DMSO-*d*₆, 400 MHz, δ; ppm) 9.81 (1H, s), 7.57 (2H, d, *J* = 7.8 Hz), 7.37–7.22 (8H, m), 7.00 (1H, t, *J* = 7.4 Hz), 4.99 (2H, s), 2.99 (2H, q, *J* = 6.5 Hz), 2.28 (2H, t, *J* = 7.4 Hz), 1.58 (2H, quintet, *J* = 7.6 Hz), 1.43 (2H, quintet, *J* = 7.1 Hz), 1.32 (2H, quintet, *J* = 7.8 Hz); MS (EI) *m/z*: 340 (M⁺).

A solution of (6-phenylcarbamoylpentyl)carbamic acid benzyl ester (1.00 g, 2.95 mmol) obtained above in MeOH (50 mL) was stirred under H₂ (atmospheric pressure) in the presence of 5% Pd/C (106 mg) at room temperature for 7 h. The catalyst was removed by filtration through Celite, and the filtrate was concentrated in vacuo. The residue was purified by silica gel flash chromatography (CHCl₃/MeOH/iPrNH₂ = 19/1/1) to give 584 mg (96%) of 58 as a white solid: ¹H NMR (DMSO-*d*₆, 400 MHz, δ; ppm) 9.83 (1H, s), 7.58 (2H, d, *J* = 7.6 Hz), 7.27 (2H, t, *J* = 7.9 Hz), 7.01 (1H, t, *J* = 7.3 Hz), 2.55 (2H, m), 2.29 (2H, t, *J* = 7.4 Hz), 1.59 (2H, quintet, *J* = 7.4 Hz), 1.37–1.30 (4H, m).

Step 3: Preparation of 6-Methanesulfonylaminohexanoic Acid Phenylamide (10). To a solution of 58 (500 mg, 2.06 mmol) obtained above in pyridine (5 mL) was added methanesulfonyl chloride (160 μL, 2.07 mmol) dropwise with cooling in an ice-water bath. The solution was stirred for 30 min at room temperature. The mixture was concentrated and diluted with AcOEt. The solution was washed with 2 N aqueous HCl, water, and brine and was dried over Na₂SO₄. Filtration and concentration in vacuo and purification by silica gel flash chromatography (*n*-hexane/AcOEt = 1/3) gave 418 mg (71%) of 10 as a crude solid. The solid was recrystallized from AcOEt to give 10 (214 mg) as colorless crystals: mp 136–137 °C; ¹H NMR (DMSO-*d*₆, 500 MHz, δ; ppm) 9.85 (1H, s), 7.58 (2H, d, *J* = 7.7 Hz), 7.28 (2H, t, *J* = 7.4 Hz), 7.01 (1H, t, *J* = 7.4 Hz), 6.93 (1H, t, *J* = 6.5 Hz), 2.92 (2H, q, *J* = 6.5 Hz), 2.87 (3H, s), 2.30 (2H, t, *J* = 7.6 Hz), 1.59 (2H, quintet, *J* = 7.6 Hz), 1.59 (2H, quintet, *J* = 7.6 Hz), 1.48 (2H, quintet, *J* =

7.4 Hz), 1.33 (2H, quintet, *J* = 7.4 Hz); MS (EI) *m/z*: 284 (M⁺); Anal. (C₁₃H₂₀N₂O₃S) C, H, N.

6-(2-Hydroxyacetyl)hexanoic Acid Phenylamide (13). To a solution of 58 (198 mg, 0.96 mmol) and glycolic acid (81 mg, 1.07 mmol) in DMF (6 mL) were added 1-ethyl-3-(3-dimethylaminopropyl)carbodiimide hydrochloride (254 mg, 1.32 mmol) and 1-hydroxy-1*H*-benzotriazole monohydrate (244 mg, 1.59 mmol), and the mixture was stirred overnight at room temperature. The reaction mixture was poured into water and extracted with AcOEt. The AcOEt layer was separated, washed with saturated aqueous NaHCO₃ and brine, and dried over Na₂SO₄. Filtration and concentration in vacuo gave 251 mg (99%) of 13 as a crude solid. The solid was recrystallized from AcOEt to give 155 mg of 13 as a colorless crystal: mp 109–113 °C; ¹H NMR (DMSO-*d*₆, 500 MHz, δ; ppm) 9.92 (1H, s), 7.79 (1H, broad s), 7.65 (2H, d, *J* = 7.6 Hz), 7.35 (2H, t, *J* = 7.9 Hz), 7.08 (1H, t, *J* = 7.3 Hz), 5.51 (1H, t, *J* = 5.8 Hz), 3.84 (2H, d, *J* = 5.8 Hz), 3.16 (2H, q, *J* = 6.8 Hz), 2.36 (2H, t, *J* = 7.5 Hz), 1.65 (2H, quintet, *J* = 7.5 Hz), 1.51 (2H, quintet, *J* = 7.3 Hz), 1.35 (2H, quintet, *J* = 7.9 Hz); MS (EI) *m/z*: 264 (M⁺); Anal. (C₁₄H₂₀N₂O₃) C, H, N.

6-(2-Aminoacetyl)hexanoic Acid Phenylamide Trifluoroacetic Acid Salt (12-TFA). Step 1: Preparation of [(5-Phenylcarbamoylpentyl)carbamoyl]methyl]carbamic Acid *tert*-Butyl Ester. This compound was prepared from 58 and *N*-(*tert*-butoxycarbonyl)glycine using the procedure described for 13 in 70% yield: ¹H NMR (CDCl₃, 400 MHz, δ; ppm) 7.53 (2H, d, *J* = 7.8 Hz), 7.34 (2H, t, *J* = 7.6 Hz), 7.10 (1H, t, *J* = 7.6 Hz), 6.14 (1H, broad s), 5.07 (1H, broad s), 3.75 (2H, d, *J* = 6 Hz), 3.30 (2H, q, *J* = 6.5 Hz), 2.37 (2H, t, *J* = 7.4 Hz), 1.76 (2H, quintet, *J* = 7.4 Hz), 1.58–1.26 (13H, m).

Step 2: Preparation of 6-(2-Aminoacetyl)hexanoic Acid Phenylamide Trifluoroacetic Acid Salt (12-TFA). To a solution of [(5-phenylcarbamoylpentyl)carbamoyl]methyl]carbamic acid *tert*-butyl ester (147 mg, 0.40 mmol) obtained above in CHCl₃ (4 mL) was added trifluoroacetic acid (1 mL), and the mixture was stirred overnight at room temperature. The reaction mixture was concentrated in vacuo, and the residue was triturated in diethyl ether to give 131 mg (84%) of 12-TFA as a white solid. The solid was recrystallized from AcOEt-MeOH to give 120 mg of 12-TFA as colorless crystals: mp 149–151 °C; ¹H NMR (DMSO-*d*₆, 500 MHz, δ; ppm) 10.00 (1H, s), 8.43 (1H, t, *J* = 5.2 Hz), 8.10 (3H, broad s), 7.71 (2H, d, *J* = 8.2 Hz), 7.41 (2H, t, *J* = 7.9 Hz), 7.14 (1H, t, *J* = 7.3 Hz), 3.25 (2H, q, *J* = 6.4 Hz), 2.43 (2H, t, *J* = 7.3 Hz), 1.72 (2H, quintet, *J* = 7.5 Hz), 1.58 (2H, quintet, *J* = 7.2 Hz), 1.44 (2H, quintet, *J* = 7.5 Hz); Anal. (C₁₄H₂₁N₃O₂·TFA·1/10H₂O) C, H, N.

6-(2-Bromoacetyl)hexanoic Acid Phenylamide (18). To a solution of 58 (70 mg, 0.340 mmol) and triethylamine (0.40 mL, 2.88 mmol) in THF (2 mL) was added a solution of bromoacetyl bromide (319 mg, 1.58 mmol) in THF (1 mL) dropwise with cooling in an ice-water bath. The mixture was stirred at room temperature for 30 min. The reaction mixture was diluted with CHCl₃, washed with aqueous saturated NaHCO₃, water, and brine, and dried over Na₂SO₄. Filtration and concentration in vacuo and purification by silica gel flash chromatography (CHCl₃/MeOH = 150/1) gave 25 mg (23%) of 18 as a white solid: ¹H NMR (CDCl₃, 400 MHz, δ; ppm) 7.52 (2H, d, *J* = 8.1 Hz), 7.32 (2H, t, *J* = 7.9 Hz), 7.19 (1H, broad s), 7.10 (1H, t, *J* = 7.6 Hz), 6.56 (1H, broad s), 3.87 (2H, s), 3.32 (2H, q, *J* = 6.6 Hz), 2.38 (2H, t, *J* = 7.3 Hz), 1.80–1.76 (2H, m), 1.63–1.59 (2H, m), 1.46–1.44 (2H, m); MS (EI) *m/z*: 326 (M⁺); Anal. (C₁₄H₁₉BrN₂O₂) C, H, N.

Thioacetic acid S-[(6-Phenylcarbamoylpentyl)carbamoyl]methyl] Ester (15). To a suspension of 18 (187 mg, 0.57 mmol) obtained above in EtOH (2 mL) was added potassium thioacetate (236 mg, 2.07 mmol), and the mixture was stirred at room temperature for 16 h. The reaction mixture was diluted with AcOEt and THF, washed with water and brine, and dried over Na₂SO₄. Filtration and concentration in vacuo and purification by silica gel flash chromatography (*n*-hexane/AcOEt = 1/1) gave 163 mg (89%) of 15 as a white solid. The solid was recrystallized from *n*-hexane-AcOEt to give 48 mg